Assignment 6-2

Name _____

Class Period _____

Date _____

1. What are **metals**?

2. How many valence electrons would be common for an atom of a **metal**?

3. How many valence electrons would be common for an atom of a **nonmetal**?

4. What happens to the outer electrons of metallic atoms when they combine chemically?

5. Define **corrosion**.

6. Give <u>two</u> examples of corrosion.

7. If an object is said to have luster, what is another word that would describe the same characteristic?

8. List five physical properties of metals.

9. Explain the term **ductile**.

10. Explain the term **malleable**.

11. In each element square on the periodic table there is a number that has a decimal point in it. What does this number represent?

12. What is the name given to the horizontal rows of elements in the periodic table?

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13. Why were elements 57 through 71 and 89 through 103 moved out of the main table and placed at the bottom?

14. Why is hydrogen placed all by itself on the periodic table?

15. Look at the diagrams in Figure 6-7. How is the electron arrangement in each element similar?

16. What are the vertical columns of elements called?

17. What do we know about the elements that are in the same vertical column?

18. Describe the left-to-right pattern that appears in each period of the periodic chart.

19. Generally speaking, how do the **physical** properties of the nonmetals compare to those of the metals?

20. What happens with the outermost energy level of nonmetals when they combine chemically?

21. Why are nonmetals with 8 electrons in the outermost energy level "special" elements?

22. What are **metalloids?**

23. List the names of the **metalloids** and their chemical symbols.

Element Name	**Symbol**

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